



Dnyansagar Coaching Classes, A'nagar

Unit Test

Std. - XI

Sub- Chemistry

(Some basic concept of chemistry)

Time - 1 hrs

Max Marks - 20

- Q.1** **Select the right option from given below -** **4**
1. The molecular weight of oxygen is -----
 - a) 16
 - b) 2
 - c) 32
 - d) 50
 2. The equivalent weight of any element is calculated by using -----
 - a) 1.008 parts by weight of hydrogen
 - b) 8 parts by weight of oxygen
 - c) 35.5 parts by weight of chlorine weight
 - d) All of these
 3. The atomacity of Argon is -----
 - a) One
 - b) Two
 - c) Three
 - d) Zero
 4. The valency of carbon is -----
 - a) 1
 - b) 2
 - c) 3
 - d) 4
- Q.2** **A. State and explain Gay-Lussac law of combining volume** **2**
- OR**
- A.** Calculate number of moles and number of molecules in 0.032 mg of methane.
- B.** Prove that Molecular weight = 2 x V.D. **2**
- C.** Explain Avogadro's law. **2**
- D.** What is equivalent weight? Explain it. **2**
- OR**
- D.** 0,082 kg of a metal gives 0.0405 kg of its chloride. Calculate equivalent mass of the metal **2**
- Q.3** **A) Define the following term. (any four)** **4**
1. Molecular weight
 2. Valency
 3. Atomic weight
 4. Vapour density
 5. Avogadro's law
 6. Gram molecular weight
- B) Answer any two** **4**
1. How equivalent weight of element is calculated by using oxidation method?
 2. What are fundamental quantities? Explain it.
 3. Explain Dulong and Petit's law? Explain it.
 4. Calculate mass of 2.8 moles of SO_2 .